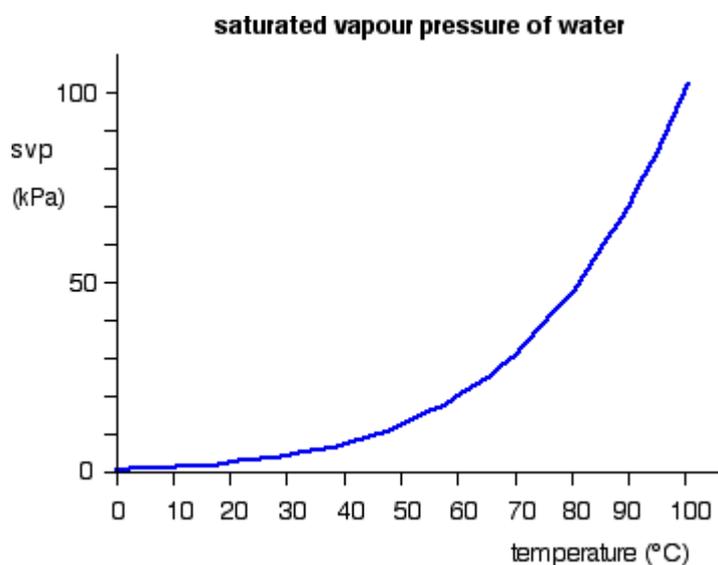


## Chemguide – questions

### SATURATED VAPOUR PRESSURE

- Explain the difference between evaporation and boiling.
  - What do you understand by the term *sublimation*?
  - By considering the evaporation of water in a sealed half-full bottle, explain what is meant by the *saturated vapour pressure* of water at a particular temperature.
  - Describe how you could measure the saturated vapour pressure of water at room temperature.
- The graph shows how the saturated vapour pressure of water varies from 0°C to 100°C.



- Explain why saturated vapour pressure increases with temperature.
- Use the graph to find an approximate boiling point of water at a pressure of 50 kPa. Explain your reasoning. (Note: If you are doing this from the graph on-screen, your answer will be very approximate! That's OK.)
- Under what circumstances will it be possible to get water to boil at temperatures greater than 100°C? Explain your reasoning.